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2.2 MASS-ACTION KINETICS 2.2.1 OVERVIEW We assume that rates can be described by $\text{Rate} = k \cdot (C)^n$ where k is the rate constant from Arrhenius' law and n is the rate law. Empirically, we say that $(C)^n = \dots$ which is typical rate law kinetics.

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J. M. Smith, Chemical Engineering Kinetics, McGraw-Hill, 1981. INSTRUCTOR BIO. Prof. Bishnupada Mandal IIT Guwahati. Dr. Bishnupada Mandal is currently a Professor and Head in the Department of Chemical Engineering at the Indian Institute of Technology Guwahati. Dr.

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